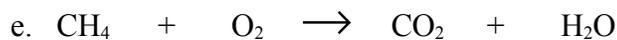
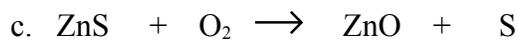
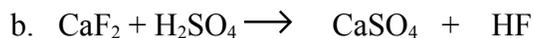


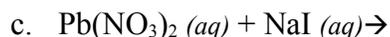
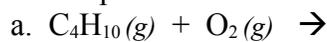
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## CHEMICAL REACTIONS REVIEW 2007

1. Explain the difference between each pair in a, b, c.
  - a.  $\text{Ca}(\text{OH})_2$  and  $\text{CaOH}_2$
  - b.  $2\text{OH}$  and  $\text{H}_2\text{O}_2$
  - c.  $\text{Ca}^{+2}$  and  $2\text{Ca}$
2. Give an example of each of the possible four states of reactants and products. What form would a precipitate be?
3. Write and balance the chemical equations for each of the following word equations:
  - a. chlorine gas and sodium bromide yield sodium chloride and bromine gas
  - b. zinc sulfide and oxygen gas become zinc oxide and sulfur
  - c. aqueous lead (II) nitrate and aqueous sodium iodide react to make solid lead (II) iodide and aqueous sodium nitrate
4. In question three, which substance was a precipitate. How can you tell?
5. Using the solubility table provided, what would be the outcome of the following reactions?
  - a.  $\text{Al}_2(\text{SO}_4)_3 + \text{Ca}(\text{OH})_2 \longrightarrow$
  - b.  $\text{KCl} + \text{AgNO}_3 \longrightarrow$
  - c.  $\text{Pb}(\text{NO}_3)_2 + \text{NaI} \longrightarrow$
6. Balance the equation and give the type of reaction
  - a.  $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$



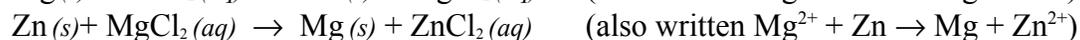
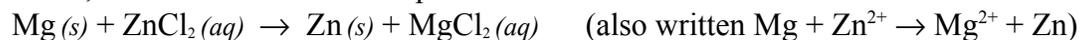
7. Predict the products of the following reactions:



Activity Series of Metals	
Ca	
Mg	
Al	
Zn	
Fe	
H	
Cu	
Ag	

Remember, that the more active metal will end up in an aqueous solution.
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8. Only one of the following reactions is likely to happen. Based on the activity series provided, which reaction will take place?



9. Describe a double replacement reaction using the letters A, B, C, D. You may need to consult your notes for assistance.

10. Oxygen gas ( $\text{O}_2$ ), hydrogen gas ( $\text{H}_2$ ), and several others are called diatomic gases. They are also called **elemental** oxygen and elemental hydrogen. What are the other five diatomic gases?