Do Now ...

Date: October 9th, 2015

Objective: Name ionic and covalent compounds.

Copy and write the name or formula:

Sulfuric Acid Hydrochloric Acid

HNO_3 Na_3PO_4

Iodine pentafluoride $Fe\ PO_4$

Monday, October 9th, 2015

Today:

W-Up, Notes & Practice: Naming and Formula Writing

HW: Go to Breslyn.org and follow the link to practice.

Unit Test: Friday

Practice ...

$CuCl_2$ copper (II) chloride

$BaCl_2$ barium chloride

$Ni(NO_3)_2$ nickel (II) nitrate

$AlBr_3$ aluminum bromide

MnO manganese (II) oxide

Naming Overview

M = metal
NM = nonmetal

$NaCl$	M + NM	Binary ionic	sodium chloride
$FeCl_3$	Transition M + NM	Binary ionic	iron (III) chloride
Na_2SO_4	M + NM + NM	Ternary Ionic (use ion table)	sodium sulfate
$FeSO_4$	Transition M + NM + NM	Ternary Ionic (use ion table)	iron (II) sulfate
H_2O	NM + NM	Molecular or Covalent	dihydrogen monoxide

More Practice ...

- phosphorus trifluoride PF_3
- Iron (II) iodide FeI_2
- chlorine monoxide ClO
- sodium ion Na^+
- lithium oxide Li_2O
- elemental nitrogen N_2
- calcium sulfate $CaSO_4$

Good to Know

Ionic = M + NM

Covalent = NM + NM

Diatomic Atoms = O_2, N_2, Br_2, Cl_2, H_2

Organic = Carbon and Hydrogen

Cation = Positive Charge

Anion = Negative Charge

Do Now ...

Date: November 10, 2015

Objective: Name and write formulas for ternary ionic compounds.

Complete:

Properties of Ionic Compounds	Properties of Covalent (Molecular)
Give an example of an ionic compound...	Example of a covalent compound ...

Tuesday, November 10, 2015

Today:

Lab: Properties of Ionic and Covalent Compounds

HW: Finish Lab Sheet

Study for Test Friday

Good to know...

For ionic we must _____.

For molecular we use _____.

For sulfates, phosphates, hydroxides, carbonates, ... we need the _____.

W-Up: Practice Naming ...

- N_2O_3 dinitrogen trioxide
- $Al_2(SO_4)_3$ aluminum sulfate
- IF_5 iodine pentafluoride
- $SiCl_4$ silicon tetrachloride
- $CuSO_4$ copper (II) sulfate
- F^- fluoride ion
- $Ca(OH)_2$ calcium hydroxide
- HCl hydrochloric acid

Do Now ...

Date: October 11th, 2015

Objective: Name and write formulas for ionic and covalent compounds.

Complete the questions on the half sheet you just picked up.

Wednesday, October 11th, 2015

Today:

W-Up, Practice, Exit Ticket

HW: Go to Breslyn.org and follow the link to practice. (Test Friday!)

Finish Exit Ticket

Good to know...

For ionic we must _____.

For molecular we use _____.

For sulfates, phosphates, hydroxides, carbonates, ... we need the _____.

Naming Overview

M = metal
NM = nonmetal

$NaCl$	M + NM	Binary ionic	sodium chloride
$FeCl_3$	Transition M + NM	Binary ionic	iron (III) chloride
Na_2SO_4	M + NM + NM	Ternary Ionic (use ion table)	sodium sulfate
$FeSO_4$	Transition M + NM + NM	Ternary Ionic (use ion table)	iron (II) sulfate
H_2O	NM + NM	Molecular or Covalent	dihydrogen monoxide

Go to www.Breslyn.org

Work on naming and writing formulas for the topics where you need to improve.

Exit ...

Complete the ½ sheet I just handed out.

Have you improved your understanding on key topics?

Do Now ...

Date: October 12th, 2015

Objective: Name ionic and covalent compounds.

Copy write the formula (use the web if you need to):

Methane: _____

Ethane: _____

Propane: _____

Thursday, October 12th, 2015

Today:

W-Up, Notes & Practice: Organic Compounds – Naming and Formulas

Study for Unit Test: Friday

More Practice Naming ...

Li_3N lithium nitride

$Mg(OH)_2$ magnesium hydroxide

PCl_3 phosphorus trichloride

sodium sulfate Na_2SO_4

$CuSO_4$ copper (II) sulfate

iron (III) chloride $FeCl_3$

N_2O dinitrogen monoxide

Do Now ...

Date: October 13th, 2015

Objective: Crush the unit test.

Quickly review your notes for the unit test today.

Also get out a blank sheet of paper.

Friday, October 13th, 2015

Today:

Unit Test